

## Chemistry of Borohydride Reduction of Iron(II) and Iron(III) Ions in Aqueous and Nonaqueous Media. Formation of Nanoscale Fe, FeB, and Fe<sub>2</sub>B Powders

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The reduction of FeBr<sub>2</sub>(aq), FeBr<sub>3</sub>(aq), FeBr<sub>2</sub>(diglyme), and FeBr<sub>3</sub>(diglyme) by NaBH<sub>4</sub> has been investigated with mechanistic and stoichiometric factors in mind. The bewildering complexity of such reactions has been partly alleviated. We find that the primary products are different depending on the reaction medium. In aqueous solution Fe<sup>2+</sup> yields mainly noncrystalline, nanoscale Fe(s), while in dry diglyme the product is nanoscale Fe<sub>2</sub>B. In the case of Fe<sup>3+</sup>(aq) the primary product again is Fe(s), but in diglyme it is FeB. An unexpected finding was that, in dry diglyme, iron borohydride coordination compounds are indefinitely stable and only yield solid Fe<sub>2</sub>B (or FeB) upon heating. These compounds are probably (L)<sub>n</sub>Fe(BH<sub>4</sub>)<sub>2</sub> and (L)<sub>n</sub>Fe(BH<sub>4</sub>)<sub>3</sub> where L = diglyme. Treatment of these complexes in-situ with water causes the precipitation of Fe<sub>2</sub>B.

### Introduction

Understanding the chemistry of BH<sub>4</sub><sup>-</sup> reduction of transition metal ions takes on increasing importance as more and more scientists use this approach to prepare "amorphous metals,"<sup>2</sup> nanocrystalline magnetic materials,<sup>3</sup> and catalysts.<sup>4</sup> It has become clear that boron is incorporated into the solid products, but the metal/boron ratios vary widely according to literature reports. This variations could be due to mixing procedures, solvent media, temperature, or other factors.

We described in recent publications the reaction steps involved in the formation of ultrafine Co<sub>2</sub>B, Co<sub>3</sub>B, metallic Co, and Co(BO<sub>2</sub>)<sub>2</sub> by the borohydride reduction of cobalt ions in aqueous and nonaqueous media.<sup>5,6</sup> As part of our continuing effort to understand and control the formation of nanoscale magnetic materials, we have studied the reduction of Fe<sup>2+</sup> and Fe<sup>3+</sup> in aqueous and nonaqueous media. Herein are reported experiments leading to different nanoscale iron compounds.

Again it has been found that the solvent medium (water or diglyme) is important in determining the final product. Likewise, subtle changes in reaction conditions can have unusual outcomes.

### Experimental Section

**General Procedure.** Reactions were carried out in a 1-L three-necked round-bottom flask attached to a vacuum line of known volume unless otherwise stated. The line was equipped with a mercury manometer, and the reaction vessel had two side arms arranged so that solutions could be added in vacuo. Acetone (Fisher Scientific) and distilled water were purged with argon for several hours prior to use. Diglyme was refluxed over sodium metal for several days and distilled under Ar. Anhydrous FeBr<sub>3</sub> (Strem Chemicals), FeBr<sub>2</sub> (Strem Chemicals) and NaBH<sub>4</sub> (Aldrich Chemicals) were stored in an inert-atmosphere box and used as received. Aqueous solutions of FeCl<sub>2</sub>·4H<sub>2</sub>O (Aldrich Chemicals) and FeCl<sub>3</sub>·6H<sub>2</sub>O (Fisher Scientific) were prepared immediately before use using prepurged water. All reactions and product isolations were carried out using standard inert-atmosphere techniques unless otherwise stated.<sup>7</sup> The amount of gas formed was determined by pressure measurement in known volume and the identity of gas evolved determined using GC.

Heat treatment of powders was carried out either in a quartz tube under Ar or in a sealed Pyrex tube under low Ar pressure. Generally, samples were brought to the desired temperature within 2 min and kept at that temperature for 2 h, after which they were allowed to rapidly cool to room temperature. For samples isolated under inert-conditions, there were no weight losses upon heat treatment. X-ray powder diffraction data were obtained on a SCINTAG 3000 machine with Cu Kα radiation. Powder diffraction studies on pyrophoric samples were carried out on samples protected with mineral oil (Nujol), and Mossbauer data were obtained on a Ranger Scientific Inc MS-1200 using a cryostat designed by Cryo Industries of America Inc. Elemental analyses were obtained from Galbraith Labs, Knoxville, TN.

**Reaction of NaBH<sub>4</sub> and FeBr<sub>2</sub> in Diglyme.** To a Schlenk tube charged with anhydrous FeBr<sub>2</sub> (0.43 g, 2.0 mmol) was added diglyme (75 mL) with stirring. An orange solution resulted after stirring for about 2 h. A second tube was charged with NaBH<sub>4</sub> (0.15 g, 4.0 mmol) in an inert-atmosphere box and 50 mL of diglyme added with stirring to dissolve the solid, yielding a colorless solution. The reaction vessel described under General Procedure was connected to a vacuum line,

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**Table 1.** Nonaqueous  $\text{FeBr}_2$  Reaction with  $\text{NaBH}_4$  under Differing Conditions<sup>a,b,c</sup>

$\text{NaBH}_4/\text{FeBr}_2$ ratio	amt of gas evolved, <sup>c</sup> mmol (expected)	reacn time, min	% yield <sup>d</sup>	tot. vol of diglyme, mL
5	7.7 (8.0)	95		125
3	8.6 (8.0)	70	80	125
2	8.2 (8.0)	60	100	125
2	8.4 (8.0)	60	81	140
2	8.2 (8.0)	90	105	200

<sup>a</sup> The  $\text{NaBH}_4$  used was dissolved in 50 mL of diglyme for all the reactions. <sup>b</sup> A total of 2.0 mmol of  $\text{FeBr}_2$  was used in all the reactions. <sup>c</sup> All decompositions were carried out thermally at about 70 °C. <sup>d</sup> Based on major product as  $\text{Fe}_2\text{B}$ .

evacuated to about  $10^{-3}$  Torr, and isolated from the pumps, and syringes were used to simultaneously transfer the diglyme solutions of  $\text{FeBr}_2$  and  $\text{NaBH}_4$  into the vessel at ambient temperature with stirring. A pale yellow solution resulted with the precipitation of a white solid ( $\text{NaBr}$ ) and minimal gas evolution. The reaction vessel was placed in a water bath and warmed to 65–70 °C. The solution became colorless (still containing the white  $\text{NaBr}$  precipitate), and vigorous gas evolution which continued for about 30 min was observed with the precipitation of a black solid. The reaction mixture was allowed to cool to room temperature. The suspension was filtered through a fine glass frit, the residue was washed with three 25 mL portions of deoxygenated water followed by two 20 mL portions of deoxygenated acetone, and the filter cake was dried in vacuo. The filter frit with the black powder was transferred into an inert-atmosphere box and the weight of the pyrophoric powder determined to be 0.12 g (0.98 mmol 98%) based on  $\text{Fe}_2\text{B}$  as the only product. An average of  $8.8 \pm 0.2$  mmol of gas was evolved during experiments carried out under these conditions.

Reactions were also carried out using different  $\text{FeBr}_2/\text{NaBH}_4$  ratios as well as different concentrations of reagents (Table 1). It should be noted that if the black suspension generated from the above reaction was filtered and the residue washed and dried under ambient conditions, a yellow-brown air-stable powder was obtained.

**Reaction of  $\text{FeBr}_2$  and  $\text{NaBH}_4$  in Diglyme in the Presence of Water.** A number of reactions were carried out as described above using a  $\text{BH}_4^-/\text{Fe}^{2+}$  ratio of 2 in which the diglyme solution of  $\text{FeBr}_2$  (0.43 g, 2.0 mmol) was treated with varying volumes of  $\text{H}_2\text{O}$  to give a desired  $\text{H}_2\text{O}/\text{Fe}^{2+}$  ratio prior to mixing with the  $\text{NaBH}_4$ -diglyme solution.

**Reaction of  $\text{FeCl}_2$  and  $\text{NaBH}_4$  in Water.** The 1000 mL reaction vessel described above was charged with  $\text{NaBH}_4$  (0.15 g, 4.0 mmol) in an inert-atmosphere box, connected to a vacuum system, evacuated, and maintained at a static pressure of about  $10^{-3}$  Torr. A freshly prepared aqueous solution of  $\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$  (0.01 M, 200 mL) was added to the vessel using four 60 mL syringes over a 45–60 s period. Immediate gas evolution was observed with the precipitation of a jellylike black solid. The reaction chamber was isolated and disconnected from the vacuum line after gas evolution had ceased (15 min). The solid was filtered off and washed with prepurged  $\text{H}_2\text{O}$  and acetone using standard inert-atmosphere techniques and the resulting powder dried in vacuo. A black, amorphous, pyrophoric powder (0.11 g) was collected in an inert-atmosphere box. Anal. Found for the crude dry powder: Fe, 82.9; B, 3.95; Fe/B ratio, 4.07. Calcd for  $\text{Fe}_2\text{B}$ : Fe, 91.2; B, 8.82. Found after thermal processing at 450 °C in a sealed Pyrex tube for 2 h: Fe, 92.2; B, 4.3; Fe/B ratio, 4.04.

The reaction described above was repeated, and the resulting black jellylike solid was isolated under ambient conditions yielding a yellow-brown air-stable powder (0.16 g).

**Decomposition of  $(\text{Diglyme})_n\text{Fe}(\text{BH}_4)_2$  by Reaction with Water.** The complex  $(\text{diglyme})_n\text{Fe}(\text{BH}_4)_2$  was prepared by reacting diglyme solutions of  $\text{FeBr}_2$  (0.43 g, 2.0 mmol) and  $\text{NaBH}_4$  (0.15, 4.0 mmol) in the 1.0-L reaction vessel as described above at ambient temperature. Addition of 1.0 mL of water (55.6 mmol) led to vigorous gas evolution with immediate precipitation of a black solid. After gas evolution had ceased (15 min), the black precipitate was isolated as described above, yielding 0.12 g of an amorphous pyrophoric powder. Analysis of samples heat processed at 450 °C using XRD and Mossbauer showed that  $\text{Fe}_2\text{B}$  (98%) was the only product. A total of 15.0 mmol of gas

was evolved. The reaction was repeated using a  $\text{BH}_4^-/\text{Fe}^{2+}$  ratio of 2 and 10 mL (556 mmol) of  $\text{H}_2\text{O}$ , yielding the same results.

**Decomposition of  $(\text{Diglyme})_n\text{Fe}(\text{BH}_4)_2$  in Air.** The  $(\text{diglyme})_n\text{Fe}(\text{BH}_4)_2$  complex prepared by reacting diglyme solutions of  $\text{FeBr}_2$  (0.43 g, 2.0 mmol) and  $\text{NaBH}_4$  (0.15 g, 4.0 mmol) in the previously described reaction vessel under inert-atmosphere conditions, was exposed to ambient laboratory conditions. A slow precipitation of a light yellowish-green solid ensued with gas evolution. After gas evolution had stopped (~2 h), the solid was filtered off, washed, and dried under ambient conditions, yielding 0.22 g of yellow-brown air-stable powder.

**Reaction of  $\text{FeBr}_3$  and  $\text{NaBH}_4$  in Diglyme.** Diglyme (50 mL) solutions of  $\text{FeBr}_3$  (0.59 g, 2.0 mmol) (red-brown) and  $\text{NaBH}_4$  (0.23 g, 6.0 mmol), prepared as described for  $\text{FeBr}_2$  above, were added to the 1.0-L reaction vessel with stirring at ambient temperature as previously described. A reddish-orange solution resulted with the precipitation of a white residue. The reddish-orange  $(\text{diglyme})_n\text{Fe}(\text{BH}_4)_3$  was stable indefinitely at ambient temperature. The reaction vessel was heated to 95–100 °C using a water bath. The solution became orange, pale yellow, and then colorless, at which time vigorous gas evolution started with the precipitation of a black solid. The reaction mixture was allowed to cool to ambient temperature and the total amount of gas evolved determined ( $9.3 \pm 0.2$  mmol). The black suspension was isolated as described for  $\text{FeBr}_2$  above, yielding 0.12 g of a black pyrophoric  $\text{Fe}_2\text{B}$  powder (86%). Anal. Found for the crude powder before heat treatment: Fe, 64.5; B, 14.1; Fe/B ratio, 0.92. Found for the powder after heat treatment: Fe, 74.7; B, 14.0; Fe/B ratio, 1.03. Isolation of the black suspension generated in the above reaction under ambient conditions yielded a brown-black air-stable powder.

**Reaction of  $\text{FeBr}_3$  and  $\text{NaBH}_4$  in Diglyme in the Presence of Water.** A series of reactions were carried out as described above using a  $\text{BH}_4^-/\text{Fe}^{3+}$  ratio of 3 in which the diglyme solution of  $\text{FeBr}_3$  (0.59 g, 2.0 mmol) was treated with varying volumes of  $\text{H}_2\text{O}$  to give a desired  $\text{H}_2\text{O}/\text{Fe}^{3+}$  ratio prior to mixing with  $\text{NaBH}_4$ -diglyme solution. Product isolation was done as previously described.

**Reaction of  $\text{FeCl}_3$  and  $\text{NaBH}_4$  in Water.** A freshly prepared 0.01 M solution of  $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$  (200 mL) was added to solid  $\text{NaBH}_4$  (0.23 g, 6.0 mmol) in a 1.0-L reaction vessel under a static vacuum, described in the general procedure, using four 60-mL syringes. Vigorous gas evolution was observed with the precipitation of a black jellylike solid. The reaction was complete in 90 s. The black suspension was isolated as described above for  $\text{FeCl}_2$ . A black amorphous, pyrophoric powder (0.11 g) was obtained. Anal. Found for the crude powder: Fe, 82.48; B, 3.58; Fe/B ratio, 4.46. After thermal processing at 450 °C in sealed Pyrex tube for 2 h analysis showed 3.70% B. If the resulting suspension from the above reaction was isolated under ambient conditions, a grayish-brown air-stable powder (0.17 g) was obtained.

**Reaction of  $(\text{Diglyme})_n\text{Fe}(\text{BH}_4)_3$  with Water.** The complex  $(\text{diglyme})_n\text{Fe}(\text{BH}_4)_3$ , prepared by reacting diglyme solutions of  $\text{FeBr}_3$  (0.59 g, 2.0 mmol) and  $\text{NaBH}_4$  (0.23 g, 6.0 mmol) as described for  $\text{FeBr}_2$ , was treated with  $\text{H}_2\text{O}$  (1.0 mL, 55.6 mmol). This resulted in vigorous gas evolution and the generation of a black suspension over 30 min. Isolation of the black solid using standard inert-atmosphere techniques yielded a noncrystalline powder (0.12 g, 98%) which upon heat processing, XRD analysis, and MS analysis indicated  $\text{Fe}_2\text{B}$ .

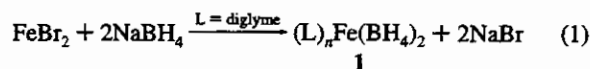
A total of 20.5 mmol of gas was evolved. When the decomposition of the precursor was carried out using 10.0 mL of  $\text{H}_2\text{O}$  (556 mmol), the reaction was complete within 1 min, yielding the same total amount of gas and solid product.

**Decomposition of  $(\text{Diglyme})_n\text{Fe}(\text{BH}_4)_3$  in Air.** Reaction of diglyme solutions of  $\text{FeBr}_3$  (0.59 g, 2.0 mmol) and  $\text{NaBH}_4$  (0.23 g, 6.0 mmol) yielded  $(\text{diglyme})_n\text{Fe}(\text{BH}_4)_3$  as previously described. Exposure of the resulting solution to ambient conditions led to gas evolution, with the reddish-orange solution becoming colorless prior to the precipitation of yellow solid. A yellow-brown air-stable powder (0.18 g) was obtained after isolation and drying in air.

## Results and Discussion

**I. Iron(II). a. In Pure Diglyme and with Subsequent Water Additions.** The interaction of iron(II) bromide and sodium borohydride in diglyme at room temperature yielded a coordination compound which we propose as 1, a pale yellow

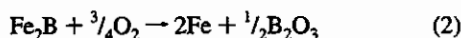
compound in solution (eq 1).



This compound is indefinitely stable at room temperature. This result is in contrast to an earlier report where  $\text{Fe}(\text{BH}_4)_2$  in diethyl ether is an unstable colorless solid that decomposes at  $-10^\circ\text{C}$ .<sup>8</sup> In our case, only upon heating to  $65\text{--}70^\circ\text{C}$  did decomposition ensue with the vigorous evolution of gas.<sup>9</sup> A noncrystalline product precipitated. After collection of this solid and drying, elemental analysis showed an Fe/B ratio of slightly less than 2. This noncrystalline material showed no XRD peaks, as expected. The Mossbauer spectra showed two overlapping broad peaks with isomer shifts of 0.20 and 0.94 mm/s with quadrupole splittings of 0.83 and 2.36 mm/s, respectively. The peaks were not well resolved even at 77 K, but these results are similar to those observed for other iron boride noncrystalline materials.<sup>10</sup>

Heat treatment of the sample yielded  $\text{Fe}_2\text{B}$  nanocrystallites along with a small amount of FeB, according to XRD and Mossbauer (Figures 1 and 2). The Mossbauer showed a hyperfine field of 237 kOe and an isomer shift of 0.08 mm/s, values that are consistent with those reported for  $\text{Fe}_2\text{B}$ .<sup>11</sup>

Exposure of the fresh noncrystalline powder to air resulted in its turning red hot and slowly smoldering. X-ray powder diffraction analysis indicated that this air-exposed sample was converted to boron oxide, iron oxides, and  $\alpha\text{-Fe}$ , probably as shown by eq 2. Evidence was also found for the presence of small amounts of nanocrystalline  $\text{Fe}_2\text{B}$  (Figure 1C,D and Figure 2C,D).



If the black suspension generated by the airless, *thermal* decomposition of 1 was isolated and dried under ambient conditions, a yellow-brown air-stable powder was obtained. Mossbauer and X-ray powder diffraction analyses of this powder are given by Figure 1E,F and Figure 2E,F and indicate that mainly iron oxides were present. This shows that slow, continuous exposure to oxygen did cause complete oxidation ( $\text{Fe}_2\text{B} + \text{O}_2 \rightarrow \text{FeO}_x + \text{B}_2\text{O}_3$ ).

The formation of 1 and its thermal decomposition to  $\text{Fe}_2\text{B}$  are independent of both the concentration of the reagents and the  $\text{BH}_4^-/\text{Fe}^{2+}$  ratio (Table 1). The yields of  $\text{Fe}_2\text{B}$  were quite high, and the amount of  $\text{H}_2$  released was very close to that

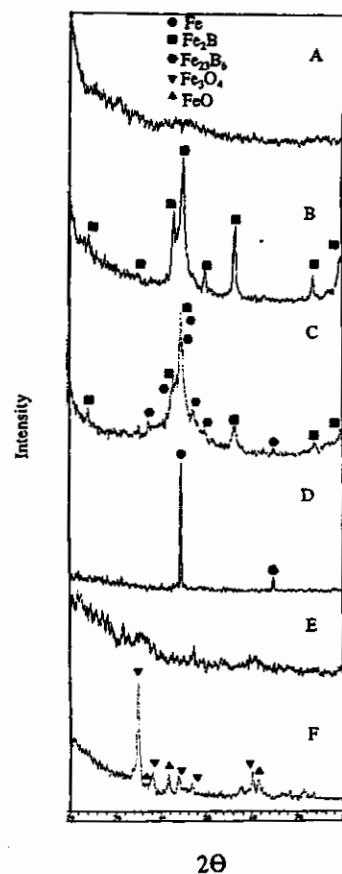


Figure 1. X-ray powder diffraction spectra of  $\text{Fe}_2\text{B}$  samples formed in nonaqueous solution: A, fresh sample protected in mineral oil; B, sample heat-treated at  $550^\circ\text{C}$ ; C, air-exposed sample; D, air-exposed sample heat-treated at  $500^\circ\text{C}$  under Ar; E, sample isolated in air, fresh; F, sample isolated in air and heat-treated at  $500^\circ\text{C}$  under Ar.

expected. These results are summarized by Scheme 1, where it is proposed that, upon heating, compound 1<sup>9</sup> decomposed with the formation of the dimer intermediate 2. The presence of this intermediate has not been confirmed; however it does have the required Fe/B ratio.

Note that  $\text{B}_2\text{H}_6$  is a primary product but that it can be catalytically decomposed in the presence of ultrafine particles of metals or metal borides<sup>6,13</sup> to give the expected amounts of  $\text{H}_2$ .

Thus, in the noncrystalline solid products both  $\text{Fe}_2\text{B}$  and B(s) should be present. After washing of the filtrate with deoxygenated water and acetone, the solid still contained a slight excess of boron. But upon heat treatment, the major nanocrystalline product was only  $\text{Fe}_2\text{B}$ . Thus, we believe most of the excess noncrystalline boron was removed by the water/acetone washing, which would be encouraged by any adventitious oxygen present as well.<sup>14</sup>

To summarize, then, these results do not *prove* that the primary product is noncrystalline  $\text{Fe}_2\text{B}$  (plus a small amount of FeB), but there is no doubt that after heat treatment nanocrystalline  $\text{Fe}_2\text{B}$  is the major product. It would seem reasonable to assume therefore, that the primary product is indeed  $\text{Fe}_2\text{B}$ , rather

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(9) Unsuccessful attempts were made to crystallize both proposed compounds  $(\text{L})_n\text{Fe}(\text{BH}_4)_2$  (1) and  $(\text{L})_n\text{Fe}(\text{BH}_4)_3$  (4). In these cases L = diglyme or triphenylphosphine. Although solids were obtained, spectral studies (IR and NMR) have still not allowed characterization. Studies are continuing. Our primary interest here is in the decomposition products.

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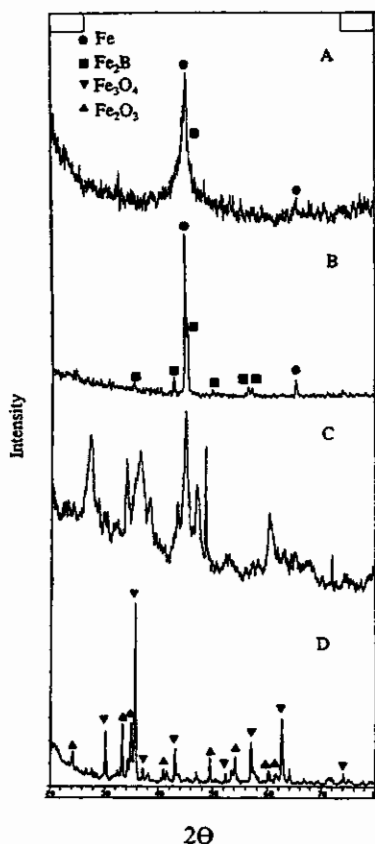
(11) Murphy, K. A.; Hershkowitz, *Phys. Rev. B* 1973, 7, 23.

(12) The unexpected phenomenon of a vigorous, exothermic oxidation leading to protected  $\alpha\text{-Fe}$  crystallites has been observed in several other systems, namely Fe-Li bimetallic particles (Glavee, G. N.; Kernizan, C. F.; Klabunde, K. J.; Sorensen, C. M.; Hadjipanayis, G. C. *Chem. Mater.* 1991, 3, 967), Fe-Mg bimetallic particles (Klabunde, K. J.; Zhang, D.; Glavee, G. N.; Sorensen, C. M.; Hadjipanayis, G. C. *Chem. Mater.* 1994, 6, 784), and  $\text{Co}_2\text{B}$  nanoparticles.<sup>5b</sup> Although such protection from oxidation originally seemed inconceivable to us as well as one of the reviewers, it is indeed what we find experimentally over and over again. This unusual phenomenon is due to the properties of the second component being both a sacrificial agent and, after oxidation, a protective agent, i.e.  $\text{Li} \rightarrow \text{Li}_2\text{O}$ ,  $\text{Mg} \rightarrow \text{MgO}$ ,  $\text{B} \rightarrow \text{B}_2\text{O}_3$ .

(13) Reaction of isolated fine powders with  $\text{BH}_3\cdot\text{THF}$  in THF in the previously described reaction setup led to the evolution of the expected amounts of  $\text{H}_2$ .

(14) Although crystalline elemental boron is chemically unreactive, nanoscale particles should be quite reactive with water (and certainly adventitious oxygen). In fact,  $\Delta H_{\text{rxn}}$  for  $2\text{B(s)} + 6\text{H}_2\text{O} \rightarrow 2\text{B(OH)}_3 + 3\text{H}_2$  is  $-445$  kJ.

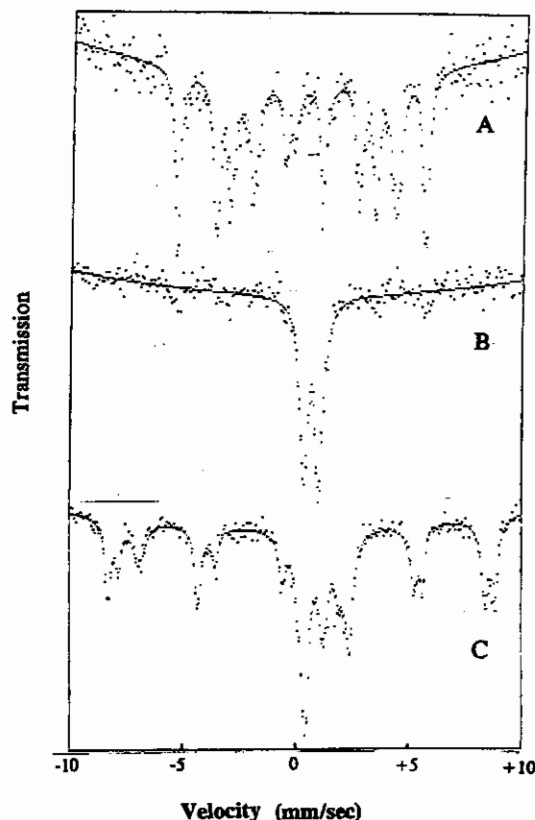




**Figure 3.** X-ray powder diffraction spectra of samples from the  $\text{Fe}^{2+}$  and borohydride reaction in aqueous solution; A, fresh sample isolated under inert-conditions; B, sample isolated under inert-conditions and heated at 500 °C under Ar; C, sample isolated under ambient conditions, fresh (a complex mixture of iron oxides and borates); D, sample isolated under ambient conditions and heated at 500 °C under Ar.

This reaction sequence attempts to account for the products and amounts formed (Table 3). For 2 mmol of  $\text{Fe}^{2+}$  14 mmol of  $\text{H}_2$  would be expected; about 15.5 was observed. This scheme attempts to show that the initial intermediate (structure 3 in Scheme 3) is apparently very susceptible to attack by a second  $\text{BH}_4^-$  by electron transfer.<sup>15</sup> Indeed, if 3 has a significant lifetime, it could dimerize, eventually resulting in  $\text{Fe}_2\text{B}$  formation (this is what is proposed in the case of  $\text{Co}_2\text{B}$  formation in aqueous media<sup>5</sup>). In fact, a small amount of  $\text{Fe}_2\text{B}$  was formed along with the major product  $\alpha\text{-Fe}$ . Thus, the fate of 3, dimerization or reduction, may be the key to understanding the disposition toward metallic iron or iron boride formation.

**II. Iron (III).** a. **In Pure Diglyme and with Subsequent Water Additions.** The addition of solutions of  $\text{FeBr}_3$ -diglyme to  $\text{NaBH}_4$ -diglyme produced a bright red color. This red solution was indefinitely stable under inert atmosphere. Upon heating of this solution to 95–100 °C, the color changed from red to orange to yellow to colorless, at which point vigorous gas evolution started and continued over a period of 45 min, and a black powder precipitated. Isolation by filtration and washing with prepurged  $\text{H}_2\text{O}$  and acetone under Ar, followed by vacuum drying, yielded a noncrystalline pyrophoric solid. Elemental analysis of this solid yielded an Fe/B ratio of 0.92. As expected, XRD of the fresh powder gave no distinguishable signal. Mossbauer at 300 and 77 K yielded a broad unresolved



**Figure 4.** Mossbauer spectra of samples from the  $\text{Fe}^{2+}$  and borohydride reaction in aqueous solution: A, sample isolated under inert-conditions and heated at 500 °C under Ar (mainly Fe(s) with some  $\text{Fe}_2\text{B}$ ); B, sample isolated under ambient conditions, fresh (mainly iron oxides); C, sample isolated under ambient conditions and heated at 500 °C under Ar (mainly iron oxides).

single line similar to that reported for noncrystalline FeB films.<sup>16</sup>

Heat treatment under Ar gave nanocrystallites of FeB (Figure 6B). Also, the analysis of this sample yielded an Fe/B ratio of 1.03. The Mossbauer spectrum of the heat-processed sample showed a complicated set of lines suggestive of two overlapping sextets with a hyperfine field of 103 kOe corresponding to the  $\alpha$ - and  $\beta$ -phases of FeB.<sup>10</sup> Thus, in the  $\text{Fe}^{3+}$  system, FeB is the primary product.

Again, as a working hypothesis, we suggest Scheme 4 as a reaction sequence leading to FeB. The formation of a stable  $(\text{L})_n\text{Fe}(\text{BH}_4)_3$  (4)<sup>9</sup> is a key component. Further work on this interesting complex is needed.

Note that for 2 mmol of  $\text{Fe}^{3+}$  there should be 12 mmol of  $\text{H}_2$  released and that the required ratio of  $\text{Fe}^{3+}/\text{BH}_4^-$  is 2/6 (rather than 2/4 for  $\text{Fe}^{2+}/\text{BH}_4^-$ ). If Scheme 4 is correct, this suggests that some of the  $\text{B}_2\text{H}_6$  formed (Scheme 4) is not completely decomposed. Control experiments have shown that nanoscale particles of  $\alpha\text{-Fe}$ , FeB,  $\text{Fe}_2\text{B}$ , Co, and  $\text{Co}_2\text{B}$  all can serve as catalysts for the decomposition of  $\text{B}_2\text{H}_6$  to boron and hydrogen.<sup>6,13</sup> However, residual  $\text{B}_2\text{H}_6$  was detected by MS in the product mix. So it is possible that the amount of  $\text{H}_2$  released by Scheme 4 was not quantitative (see Table 4).

When the fresh noncrystalline FeB solid product was quickly exposed to air, it spontaneously began to smolder with a red-hot heat. XRD and Mossbauer analysis of crystalline products showed the presence of  $\alpha\text{-Fe}$ ,  $\text{Fe}_2\text{B}$ ,  $\text{Fe}_{23}\text{B}_6$ , and FeB. Presum-

(15)  $\text{BH}_4^-$  can readily serve as both a hydride and electron transfer agent in organic and inorganic reactions. See: Cotton, F. A.; Wilkinson, G. *Advanced Inorganic Chemistry*, 5th ed.; New York, 1988; pp 190–192.

(16) (a) Blum, N. A.; Moorjani, K.; Poehler, J. O.; Satkiewicz, F. G. *J. Appl. Phys.* 1982, 53, 2074. (b) Blum, N. A.; Moorjani, K.; Poehler, J. O.; Satkiewicz, F. G. *J. Appl. Phys.* 1981, 52, 1808.

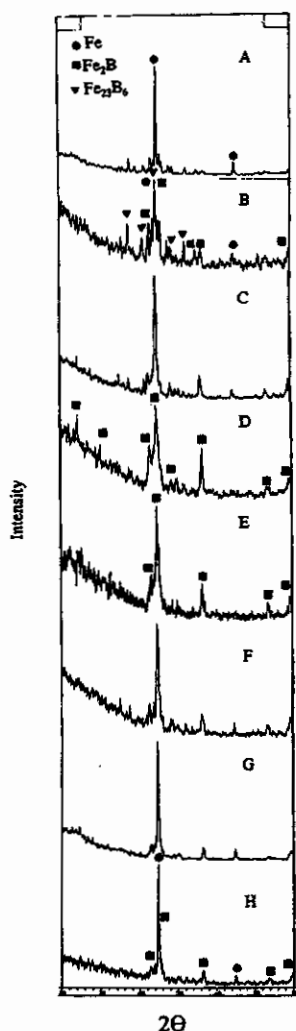
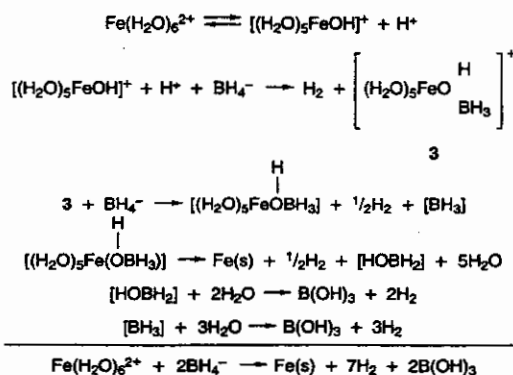


Figure 5. X-ray powder diffraction spectra showing the changes in the product for the nonaqueous  $\text{FeBr}_2$  and  $\text{NaBH}_4$  reaction in the presence of  $\text{H}_2\text{O}$ ; A, 0.22 mL of  $\text{H}_2\text{O}$  added; B, 1.0 mL of  $\text{H}_2\text{O}$  added; C, 5 mL of  $\text{H}_2\text{O}$  added; D, 10 mL of  $\text{H}_2\text{O}$  added; E, 15 mL of  $\text{H}_2\text{O}$  added; F, 30 mL of  $\text{H}_2\text{O}$  added; G, 50 mL of  $\text{H}_2\text{O}$  added; H, 100 mL of  $\text{H}_2\text{O}$  added. Samples were heat-treated at  $400^\circ\text{C}$ .

### Scheme 3. Iron(II) Bromide (Aqueous)



ably, noncrystalline  $\text{B}_2\text{O}_3$  was also formed. This example shows again that oxidation reactions can be deceiving and often allow metal and metal boride particles to be formed and protected from further oxidation by a  $\text{B}_2\text{O}_3$  coating.<sup>5</sup> The formation of  $\text{B}_2\text{O}_3$  from FeB is extremely exothermic, and so the oxidation-process can generate enough heat to encapsulate nonoxidized

Table 3. Aqueous Sodium Borohydride Reduction of Iron(II) and Iron(III) Ions

$\text{NaBH}_4/\text{iron}$ ratio <sup>c</sup>	amt of gas evolved, mmol	pH of filtrate <sup>f</sup>	reacn time, min	% yield <sup>g</sup>
2 <sup>a</sup>	15.5	6.8	15	98
3 <sup>a</sup>	24.6	9.0	60	94
2 <sup>a,d</sup>	15.4	6.2	15	
2 <sup>a,e</sup>	15.0	7.9	60	90
3 <sup>b</sup>	23.0	5.6	1.5	98
3 <sup>b,d</sup>	23.0	7.0	1.5	
3 <sup>b,d,e</sup>	22.0	5.2	5.0	

<sup>a</sup> Iron(II) reduction reaction. <sup>b</sup> Iron(III) reduction reaction. <sup>c</sup> All reactions carried out by adding 200 mL of 0.01 M iron salt solution to the appropriate weight of  $\text{NaBH}_4$  solid. <sup>d</sup> Sample isolated under ambient conditions. <sup>e</sup> The 200 mL iron salt solution was added to the solid  $\text{NaBH}_4$  in two 100 mL portions, waiting for all gases to be evolved before addition of the second portion. <sup>f</sup> pHs of starting iron(II) and iron(III) solutions were 3.6 and 2.2, respectively. <sup>g</sup> Based on major product as  $\alpha$ -Fe, except in cases of isolation under ambient conditions.

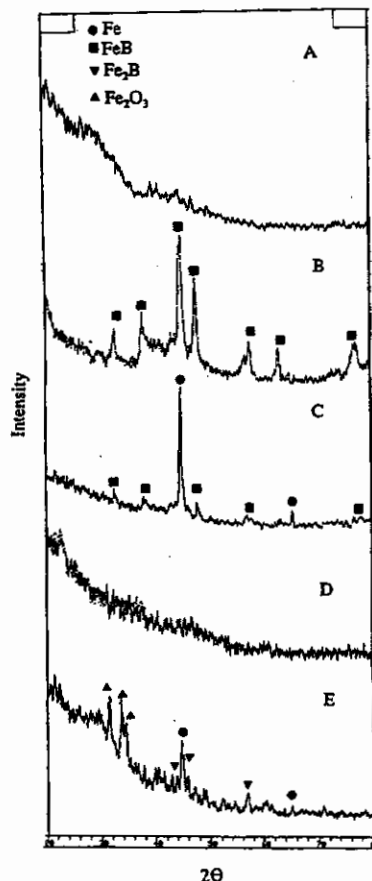
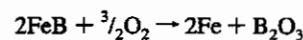


Figure 6. X-ray powder diffraction spectra of FeB samples generated in nonaqueous solution: A, fresh sample protected in mineral oil; B, sample heat-treated at  $400^\circ\text{C}$  under Ar; C, sample exposed to air; D, sample isolated under ambient conditions; E, sample isolated under ambient conditions and heat-treated at  $500^\circ\text{C}$  under Ar.

materials.



$$\Delta H_{\text{reacn}}^\circ \approx -2400 \text{ kJ}$$

In another set of experiments, the workup and isolation of the fresh FeB product were carried out in air. A brown-black air-stable powder was collected. Upon heat treatment under



